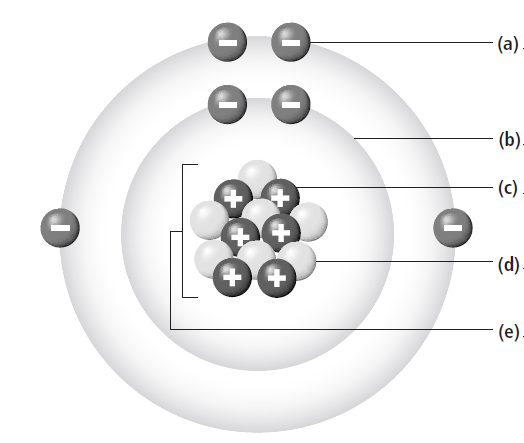
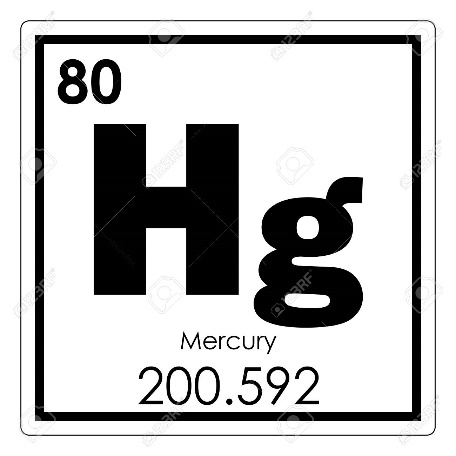
Sci 9 **Atomic Theory Practice Quiz**

**Atomic History: Who said what?**

1. Atoms cannot be created, destroyed, or divided into smaller particles.
2. Electrons occupy specific energy levels or shells.
3. Most of the mass of the atom is in the tiny, dense, positively charged nucleus.
4. Most of the atom is empty space.
5. All atoms of the same element are identical.
6. Atoms contain negatively charged particles.
7. Nucleus contains positive particles called protons and neutral particles called neutrons.
8. Electrons move around a central nucleus

**Label the parts of the atom:**

How many protons, neutrons, and electrons does one atom of mercury (below) have?



1. What is the definition of an atom? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
2. What two things does a Bohr model show? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
3. What subatomic particles make up the atom’s mass?
4. What is the charge of an atom?
5. What is the charge of an atom’s nucleus?
6. What is the definition of a composite particle?
7. Give two examples of composite particles.
8. Give three examples of elementary particles.