

Writing Chemical Names and Formulas of Ionic Compounds

Use with textbook pages 154-161.

You can use the periodic table on page 106 to help you answer these questions.

1. Define the following terms:

a) ion

b) ionic bond

c) binary ionic compound

2. Write the formulas for the ionic compounds formed from the following positive and negative ions. Then name the binary ionic compounds.

	Ions	Chemical Formula	Compound Name
a)	Li^+ F^-		
b)	Be^{2+} Br^-		
c)	Na^+ O^{2-}		
d)	Sc^{3+} I^-		
e)	Mg^{2+} S^{2-}		
f)	K^+ P^{3-}		
g)	Ca^{2+} N^{3-}		
h)	Al^{3+} Br^-		
i)	Ba^{2+} Se^{2-}		
j)	Zn^{2+} O^{2-}		

3. Write the names of the following ionic compounds.

- a) RbI _____
- b) Cs₂O _____
- c) SrS _____
- d) BaCl₂ _____
- e) InP _____
- f) Li₃N _____
- g) MgBr₂ _____
- h) CaO _____
- i) KCl _____
- j) ZrSe₂ _____

4. Write the chemical formulas for the following ionic compounds.

- a) cesium fluoride _____
- b) beryllium phosphide _____
- c) aluminum oxide _____
- d) strontium bromide _____
- e) gallium nitride _____
- f) zinc sulfide _____
- g) rubidium iodide _____
- h) silver chloride _____
- i) cadmium oxide _____