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2. Alkali Metals are in the first column (Group 1)

Soft, low density, can be cut with a knife

React with air and water

Eg. sodium, potassium (used in fertilizer), rubidium, cesium

3. Uses of Alkali metals:

Lithium used in batteries, sodium to combine with other compounds, potassium in fertilizer

4. uses of Alkaline Earth Metals:

Magnesium used in flash powder and fire works

Magnesium and beryllium are imporatn for many industrial alloys

Magnesium added to other metals for strength but maintain lightness

5. Noble gas are unreactive

6. Uses of noble gases: Neon signs, argon in light bulbs, helium used in lab due to its ability to maintain low temperatures

7. Halogens have bright gases

8. Chlorine – component of plastics and bleach

9. hydrogen – family of one – can be have as a highly reactive metal or a hight reactive non-metal because it can either give up or acquire an electron forming either positive or negative ions.